

10/2/2009

**Home assignment:**

**Chapter 9.** – Read Ch.9 pp.329-372 (aim to complete by Wed, Oct.14)

**Recommended problems:** on pp. 366-371: 9.XX, where XX is:

1, 6, 9, 17, 27,  
29, 30,33,35,36,  
41, 35,46,49,52,  
55, 57, 51, 65, 69, 75,  
77, 81, 83, 85, 87,89,  
93,97,98,105,107  
harder/cumulative-skill problems:127, 129

(Remember, the blue ones have solutions in the end of the book – work out your own solution first, THEN check the book solution. Be sure you understand it.)

**Some of the important terms** (define/explain in writing):

Ionic bond compared to covalent bond; lattice energy for ionic solid, bonding and non-Lewis formulas: bonding and non-bonding (lone) electron pairs; single, double and triple bonds; electronegativity and polar covalent bonds; delocalized bonding and resonance Lewis formulas; formal charge in Lewis formulas, bond order and bond length.

**developing skills:**

1. writing electron configurations of ions and showing ion formation with Lewis dot structures
2. comparing relative ionic sizes based on periodic trends
3. interpreting covalent bond polarity based on the electronegativities of individual atoms
4. drawing Lewis structures based on molecular formula or formula of an ion.
5. writing resonance Lewis formula
6. relating bond order and bond length