

out of 10

SHOW ALL THE explanations asked .

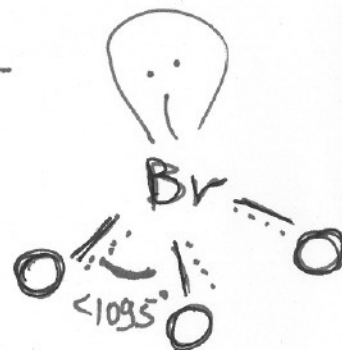
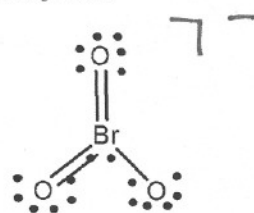
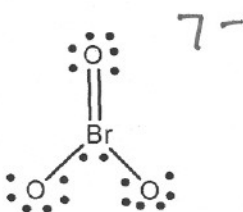
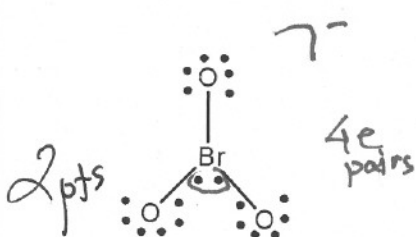
Put letter answer in the space provided next to each question.

- C 1. Three of the four following species: XeF_4 , SiH_4 , ClO_4^- and NH_4^+ would be expected to have the same molecular geometry. What is this geometry?
 A. square planar B. pyramidal C. tetrahedral D. see-saw
 Sketch molecular geometry and name it for each compound listed above.

2 pts

square planar

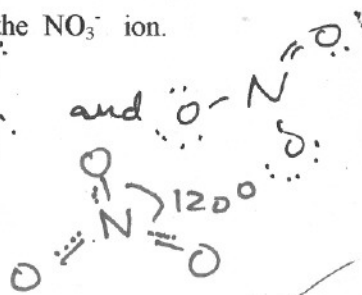
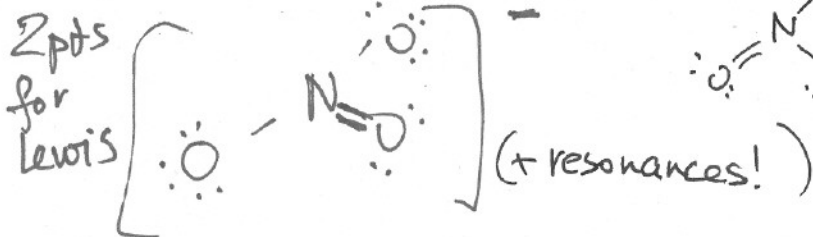
- C 2. The students in class proposed three Lewis structures for the BrO_3^- ion :



All three structures would result in the SAME molecular geometry. What is this geometry?

- A. trigonal planar B. tetrahedral C. Pyramidal D. T-shape E. square planar

- D 3. Sketch the Lewis structure and molecular geometry for the NO_3^- ion.

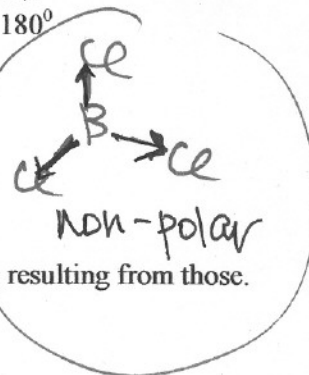
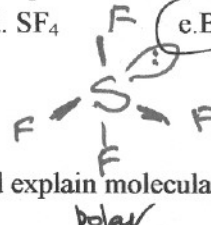
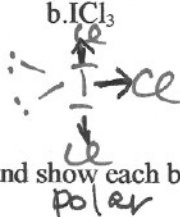
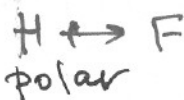


What is the closest value of predicted ONO angle (put letter answer next to the question above)?

- a. 60° b. 90° c. 109° d. 120° e. 180°

- E 4. Which of the following has polar bonds but is a non-polar molecule?

- a. HF b. ICl_3 c. NE_3 d. SF_4 e. BCl_3



Sketch this molecule shape and show each bond dipole vectors and explain molecular dipole resulting from those.

DIFFERENT QUESTION!!!!

1 pt actual drawing